PETA KONSEP
CONCEPT MAP

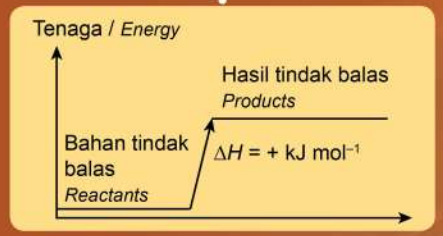
Termokimia
Thermochemistry

Perubahan haba
Heat change
$Q = mc\theta$ J

Tindak balas endotermik
Endothermic reaction
ΔH = positif
positive

Pemutusan ikatan dalam bahan tindak balas menyerap haba
Breaking of bonds in reactants absorbs heat

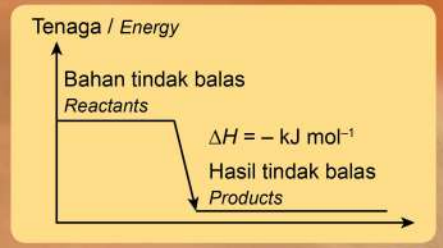
Aplikasi
Application
Pek sejuk, minuman sejuk kendiri, serbat
Cold pack, self-cooling beverages, sorbet



Tindak balas eksotermik
Exothermic reaction
ΔH = negatif
negative

Pembentukan ikatan dalam hasil tindak balas membebaskan haba
Formation of bonds in products releases heat

Aplikasi
Application
Pek panas, hot can, pembakaran bahan api
Hot pack, hot can, combustion of fuels



Haba tindak balas
Heat of reaction
ΔH kJ mol^{-1}

Perubahan tenaga haba yang berlaku apabila 1 mol sesuatu bahan tindak balas bertindak balas dengan lengkap atau apabila 1 mol sesuatu hasil tindak balas terbentuk pada keadaan piawai.
Heat change that occurs when the 1 mole of reactants react completely or 1 mole of products form under standard conditions.

Haba pemendakan
Heat of precipitation
$\Delta H = -$kJ mol^{-1}
$\Delta H = +$kJ mol^{-1}

Haba penyesaran
Heat of displacement
$\Delta H = -$kJ mol^{-1}

Haba peneutralan
Heat of neutralisation
$\Delta H = -$kJ mol^{-1}

Haba pembakaran
Heat of combustion
$\Delta H = -$kJ mol^{-1}

Nilai bahan api
Fuel value

© Penerbitan Pelangi Sdn. Bhd.